Bond breaking and making: the energy account

Teacher / technician

YEAR 11 / LESSON 85 / CHEMISTRY ONLY

Staff preparation and answer guidance. Keep separate from student sheets. Written and model-based lesson; no laboratory procedure is implied.

Coverage and delivery

Chemistry only: 3.6C. Mean bond-energy bookkeeping, not an experiment or reaction mechanism.

Allow about 60 minutes. Follow the student sequence and use the answers to check reasoning.

Requisition • class of 30, working in 15 pairs

Item	Per pair / setup	Class total
Booklet / writing tools	1 per learner	30 sets
Model kit / periodic table	1 set per pair	15 sets

Immediate safety

Desk activity only; keep small model pieces away from mouths.

Preparation and delivery

Read alongside the student method

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1. Provide 30 booklets and 15 model kits or reusable atom counters; no chemicals are needed.

2. Use 10 minutes retrieval, 25 independent/pair tasks, 15 explanation and 10 corrections. Keep first attempts visible.

3. Use arrow direction and sign consistently; do not describe a bond itself as a container that spills energy on breaking.

4. Keep this Chemistry-only; shared-core learners are not assigned the bond-energy extension.

5. Check all carbon valencies, formula subscripts and equation coefficients. Label any supplied numbers as practice data, not class observations.

Close of lesson

• Return shared equipment and retain corrected reasoning.

Check before teaching

Check the cards, reference data and quantities against this requisition. Keep the staff answer guide separate until the checking stage. Supplied numerical practice datasets are labelled examples; they are not class measurements.

Answers and teaching notes

Expected explanations • do not give as experimental observations

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- Endothermic: energy is required to separate bonded atoms.
- Exothermic: energy is released when the bond forms.
- 500−650=−150 kJ mol^{-1}; exothermic.
- 700−550=+150 kJ mol^{-1}; endothermic.
- Activation energy is needed to start reaction; the overall balance can still release more energy than it absorbs.
- Breaking consumes energy; net release occurs when forming product bonds releases more than breaking reactant bonds requires.

References and provenance

Pearson Chemistry specification

Prepared 22 November 2026. Original wording, tables and questions. Sources inform the concepts and technique; the lesson scope and classroom quantities are explicitly described here. No laboratory trial or final school risk approval is claimed.